The kinetic constants of the two enzymes. These results confirmed two earlier studies on intact erythrocytes, which refuted the notion that glutathione peroxidase is the primary enzyme for removal of hydrogen peroxide within erythrocytes. The present findings differ from the results with intact cells, however, in showing that glutathione peroxidase accounts for even less than 50% of the removal of hydrogen peroxide. A means is proposed for calculating the relative contribution of glutathione peroxidase and catalase in other cells and species. The present results raise the possibility that the major function of glutathione peroxidase may be the disposal of organic peroxides rather than the removal of hydrogen peroxide.

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order kinetic constant (sec⁻¹) for disappearance of H₂O₂, as determined at a wavelength of 240 nm with a recording spectrophotometer. When multiplied by the concentration of H₂O₂, k gives the rate at which H₂O₂ is used at that concentration. Catalase activity per unit volume (k/mL) of a solution was based on the first-order kinetic constant observed when an aliquot of a solution was added to an assay mixture. The kinetic constant was then multiplied by the dilution factor (assay volume/aliquot volume). A unit (U) of activity of glutathione reductase and glutathione peroxidase represents an enzymatic rate of one μmole of NADPH oxidized per minute. Assays of glutathione peroxidase were performed at 37°C by the method of Beutler, as follows, except that the reaction was performed at pH 7.4 in KRT buffer. The reaction mixture had a final volume of 1 mL and contained, at the following final concentrations: GSH 2 mmol/L, NADPH 200 μmol/L, t-butyl hydroperoxide 70 μmol/L, and yeast glutathione reductase (2 U/mL of reaction mixture). The reaction was followed from the rate of disappearance of NADPH, as determined at 340 nm with a recording spectrophotometer. The presence of a relative excess of glutathione reductase served to keep the GSH at a constant concentration of 2 mmol/L. Because catalase has almost no affinity for t-butyl hydroperoxide, this substrate provides a reliable measurement of glutathione peroxidase in the presence of erythrocytic catalase. After collection, partial purification and concentration of the separate fractions (catalase, glutathione peroxidase, and hemoglobin), the three preparations were mixed in a reaction mixture in the same proportions as were present in the crude hemolysates. The final concentrations of catalase purified from human erythrocytes varied between 0.16 k and 0.60 WmL, in which k represents the first-order rate constant (sec⁻¹) of catalase. Final concentrations for human purified glutathione peroxidase were between 0.016 and 0.06 U/mL. The final concentration of human hemoglobin, when present, was between 10 and 20 μmol/L according to the concentration of the two antioxidant enzymes present in the reaction mixtures. The reaction was then exposed to H₂O₂, which was generated at a constant rate by glucose and glucose oxidase. The reaction mixture had a final volume of 1.0 mL and contained: GSH 2 μmol, NADPH 2 μmol, glucose-6-phosphate 1.0 μmole, glutathione reductase (yeast) 2 U, glucose-6-phosphate dehydrogenase (yeast) 10 μg, glucose 5.0 μmol, and glucose oxidase (5 picomoles).

Some experiments were performed with commercially available, highly purified catalase and glutathione peroxidase from human erythrocytes. Before use, the enzymes were thoroughly washed on ultrafiltration cones to remove all preservatives, then resuspended in KRT. Activities of these enzymes were measured as for hemolysates just before they were added to incubation mixtures. Incubations were in Pyrex tubes at 37°C for 20 to 30 minutes in a metabolic shaker. The reaction was stopped at the end of the incubation by heating of the tubes at 100°C for 2 minutes. Supernatant fluid was removed after the tubes were centrifuged at 2000g for 2 minutes, then kept on ice before determination of the amount of 6-phosphogluconate (6PG) that had been generated during the incubation. Concentrations of 6PG were measured with a spectrophotometer at 340 nm in a 0.5 mL reaction mixture containing 0.1 mol/L Tris-HCl, pH 8.0, 0.05 mol/L MgCl₂, 1 mmol/L NADP⁺, 5 μg yeast 6PGD, and 50 μL of the incubation mixture. The net increase in NADPH corresponded to the amount of 6PG generated during the exposure of catalase and glutathione peroxidase to the H₂O₂ flow. All incubations and determinations were in duplicate.

RESULTS

At a GSH concentration of 2 mmol/L, the Michaelis constant (Km) for H₂O₂ of purified, human erythrocytic glutathione peroxidase was found to be 28 μmol/L and the enzyme used H₂O₂ at a rate that was 2.5 times the rate at which it used t-butyl hydroperoxide. The erythrocytes of nine normal Caucasian subjects contained 28.7 ± 6.4 (mean ± standard deviation [SD]) U of glutathione peroxidase per gram of hemoglobin, and the catalase activity was 312 ± 44 k per gram of hemoglobin. Whether human glutathione peroxidase and catalase were purchased or partially purified, cost and practical limitations prohibited incubations of the two enzymes at intracellular concentrations, especially given the need for a relative excess of reagent glutathione reductase. For the in vitro metabolic studies, these two enzymes were kept at concentrations proportional to the erythrocytic values, but at levels well below the erythrocytic concentrations. At the intracellular ratio of catalase to glutathione peroxidase and at a constant rate of generation of H₂O₂, the rate of generation of 6PG remained almost the same over a range of concentrations of the two enzymes (Table 1).

The rate of formation of 6PG was 1.33 ± 0.33 nmol mL⁻¹ min⁻¹ at a concentration of 0.6 μmol/L of catalase and 0.06 μmol/L of glutathione peroxidase in the presence of an excess of G6PD and glutathione reductase and at a constant flow of H₂O₂ of 12 nmol mL⁻¹ min⁻¹ (Fig 2). In the presence of glutathione peroxidase only, the generation of 6PG was 9.73 ± 1.47 nmol mL⁻¹ min⁻¹, in agreement with the rate expected from the fact that one mole of 6PG is formed for each mole of H₂O₂ reduced. Exposure of catalase to the same flow of H₂O₂, in the absence of glutathione peroxidase, led to the generation of 6PG at a rate of 0.21 ± 0.03 nmol mL⁻¹ min⁻¹ (Fig 2). In the presence of both enzymes in the

Table 1. Rate of 6PG Generation During Exposure of Catalase and Glutathione Peroxidase, at Different Concentrations, to H₂O₂ at the Same Rate of H₂O₂ Flow (5.3 nmol mL⁻¹ min⁻¹)

<table>
<thead>
<tr>
<th>Catalase (μmol/L)</th>
<th>Glutathione Peroxidase (μmol/L)</th>
<th>6PG (nmol mL⁻¹ min⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.15</td>
<td>0.015</td>
<td>1.62</td>
</tr>
<tr>
<td>0.30</td>
<td>0.030</td>
<td>1.34</td>
</tr>
<tr>
<td>0.60</td>
<td>0.060</td>
<td>1.36</td>
</tr>
<tr>
<td>1.00</td>
<td>0.100</td>
<td>1.33</td>
</tr>
</tbody>
</table>

A system of regeneration of GSH and NADPH was present (see Materials and Methods).
CATALASE AND HUMAN ERYTHROCYTES

Fig 2. Generation of 6PG over 30-minute during exposure of catalase, glutathione peroxidase, and both enzymes, to a H2O2 flow of 12.0 nmol mL\(^{-1}\) min\(^{-1}\) in the presence of constant concentration of GSH (2 mmol/L) and NADPH (2 μmol/L). Values represent the mean ± SD of seven experiments.

erythrocytic proportion, the glutathione peroxidase, therefore, handled only 15% of the generated H2O2 (Fig 2). Exposure of the same reaction mixture to different rates of H2O2 generation elicited proportional amount of 6PG for all three enzyme combinations (Fig 3). The higher values of 6PG observed at zero flow of H2O2 may be related to autooxidation of reduced glutathione that is dependent on temperature and pH. During the incubation, which never exceeded 30 minutes, the rate of generation of 6PG remained constant. Under the conditions specified above, the rate of formation of 6PG was not greatly altered by the addition to the catalase/glutathione peroxidase system of spin-trapping agents or such other enzymes as NADPH diaphorase, superoxide dismutase (Fig 2 and Table 2) or glutathione-S-transferase (not shown). A small, but consistent, stimulation of the rate of 6PG production was seen when purified human hemoglobin was present (Table 2). Several incubation mixtures, containing decreasing amount of catalase in the presence of a constant activity of glutathione peroxidase, were prepared and were exposed to the same flow of H2O2 for 30 minutes. At the catalase/glutathione peroxidase ratio observed in normal erythrocytes, the formation of 6PG was about one-fifth to one-tenth of that expected if the glutathione peroxidase were fully active (far right of Fig 4).

Expected rates of 6PG formation were assumed to be the rate at which glutathione peroxidase was functioning in a hypothetical reaction mixture containing glutathione peroxidase, catalase, and GSH (2 mmol/L) at known concentrations and H2O2, at a known and constant rate of generation. The maximal velocity of the glutathione peroxidase was taken as the activity with H2O2 (100 μmol/L) as the substrate. The Km of the glutathione peroxidase was considered to be 28 μmol/L. An assumption that the reaction had reached a steady state (constant H2O2 concentration) allowed the concentration of H2O2 to be calculated. This value allowed calculation, in turn, of the rate at which the glutathione peroxidase was functioning. The observed rate of 6PG formation was similar to the expected rate, and all rates followed a relatively smooth curve upward at progressively lower concentrations of catalase (Fig 4). With the addition of purified hemoglobin, the curve was shifted upward slightly, but remained concave. Curves of similar shape were observed when commercially prepared enzymes were used (Fig 5).

DISCUSSION

Until mid-century, catalase was regarded as the enzyme responsible for disposing of H2O2. In 1957, Mills discovered glutathione peroxidase and demonstrated that it used GSH

Table 2. Effect of Different Agents on the Rate of Production of 6PG During Exposure of Catalase and Glutathione Peroxidase to H2O2 Flows (7.2 nmol mL\(^{-1}\) min\(^{-1}\)) in the Presence of a System for Regeneration of NADPH and GSH

<table>
<thead>
<tr>
<th>Addition (final concentration)</th>
<th>6PG Generation (% of baseline values)</th>
</tr>
</thead>
<tbody>
<tr>
<td>None</td>
<td>100</td>
</tr>
<tr>
<td>NADPH diaphorase (9 μmol/L)</td>
<td>125</td>
</tr>
<tr>
<td>Superoxide dismutase (5 μg/mL)</td>
<td>105</td>
</tr>
<tr>
<td>Hemoglobin (10 nmoles/mL)</td>
<td>116</td>
</tr>
<tr>
<td>Hemoglobin plus superoxide dismutase</td>
<td>106</td>
</tr>
<tr>
<td>Methemoglobin (10 nmoles/mL)</td>
<td>114</td>
</tr>
<tr>
<td>N-tert-butyl-a-phenylnitrone (1 mmol/L)</td>
<td>100</td>
</tr>
<tr>
<td>Hemoglobin plus N-tert-butyl-a-phenylnitrone (1 mmol/L)</td>
<td>126</td>
</tr>
</tbody>
</table>

Values are expressed as a percentage of the rate of 6PG generation without an added agent.
remain active in the constant presence of H$_2$O$_2$. In two thione reductase, using NADPH. A prevailing view that concentration of hemoglobin was 10 pmol/L.

Previous studies reported in this journal, we used intact human erythrocytes to estimate the relative contributions of catalase and glutathione peroxidase in the removal of H$_2$O$_2$. This view derived from studies on erythrocytes with a reduced capability to generate adequate amounts of NADPH, of which G6PD deficiency is the paradigmatic example. Catalase was regarded as a fossil enzyme. That catalase, itself, requires NADPH was suggested, however, by the early studies of Eaton et al and was proved by the more recent discovery that mammalian catalase has tightly bound NADPH and uses NADPH to remain active in the constant presence of H$_2$O$_2$. In two previous studies reported in this journal, we used intact human erythrocytes to estimate the relative contributions of catalase and glutathione peroxidase in the removal of H$_2$O$_2$. Normal cells were compared with those with a genetic deficiency of catalase or with varying degrees of aminotriazole-induced inactivation of catalase. We concluded that catalase accounts for the disposal of about half of the H$_2$O$_2$ that might be generated at physiologically and pathologically realistic rates. Glutathione peroxidase was presumed to account for the other half. Anomalies in the kinetics of H$_2$O$_2$ removal in the second study, however, made it difficult to attribute half of the H$_2$O$_2$ disposal entirely to glutathione peroxidase. Moreover, glutathione peroxidase was assumed in those two studies to be the only mechanism by which the removal of H$_2$O$_2$ was accompanied by the oxidation of NADPH or GSH.

The present cell-free system consisted of a mixture of the enzymes involved in the disposal of H$_2$O$_2$. This system allowed the concentration of individual components to be changed so as to provide better insight into the H$_2$O$_2$-removing role of each component than could be gained with studies of intact cells. Both the observed and predicted contributions by glutathione peroxidase were much less than those of catalase. Calculations giving the expected rate of glutathione peroxidase gave also an estimate of the steady-state concentration of H$_2$O$_2$. Under the conditions of each of the tables and figures in this report, the concentrations of H$_2$O$_2$ were well below 28 μmol/L, the Km for H$_2$O$_2$ of glutathione peroxidase at a GSH concentration of 2 mmol/L. At such relatively low concentrations of H$_2$O$_2$, the activities of catalase and glutathione peroxidase would each be proportional to the H$_2$O$_2$ concentration, and glutathione peroxidase would remove H$_2$O$_2$ at a rate that is Vmax/kKm times the rate at which H$_2$O$_2$ would be removed by catalase. Vmax and Km are the constants of glutathione peroxidase at a specified concentration of GSH, and k is the first-order rate constant of catalase. For any given ratio of glutathione peroxidase to catalase activity, therefore, the fraction of total H$_2$O$_2$ that is removed by glutathione peroxidase should remain the same over a range of concentrations of the two enzymes or of rates of generation of H$_2$O$_2$. That expectation is confirmed by the observations of Table 1 and Fig 3. This approach provides an estimate of the maximum contribution of glutathione peroxidase, relative to that of catalase, even though the concentration of H$_2$O$_2$ may be unknown. Activity of catalase is simply proportional to the H$_2$O$_2$ concentration, whereas activity of glutathione peroxidase follows Michaelis-Menten relationships and begins to level off as the concentration of H$_2$O$_2$ gets above the Km of glutathione peroxidase. We consider the generation rates and concentrations of H$_2$O$_2$ in this and our earlier studies to be

![Graph showing the rates of formation of 6PG in the presence of increasing amounts of catalase and at fixed concentrations of glutathione peroxidase](image)

**Fig 4.** Rates of formation of 6PG in the presence of increasing amounts of catalase and at fixed concentrations of glutathione peroxidase (0.06 U/mL), NADPH (2 μmol/L), and GSH (2 mmol/L). The rate of generation of H$_2$O$_2$ was constant (12.0 nmol mL$^{-1}$ min$^{-1}$). The concentration of hemoglobin was 10 μmol/L.

![Graph showing the relative 6PG generation](image)

**Fig 5.** Relative rates of production of 6PG during incubation of human glutathione peroxidase and catalase (from commercial enzymes) to H$_2$O$_2$ generated at a constant rate (12 nmol mL$^{-1}$ min$^{-1}$) by the glucose plus glucose oxidase system in the presence of NADPH (2 μmol/L), GSH (1 mmol/L), GSH (2 mmol/L), yeast G6PD (10 μg/mL), and yeast glutathione reductase (2 U/mL). Dotted line represents the relative hexose monophosphate shunt activity of previous experiments performed in intact erythrocytes. Bars represent the SD of five experiments.
toxicologically realistic. Those who believe that the rates and concentration are even higher must contend with the fact that the activity of glutathione peroxidase, relative to that of catalase, would then be even lower.

Estimates of the expected rate of glutathione peroxidase at low concentrations of \( \text{H}_2\text{O}_2 \) would be sensitive to errors in the determination of the \( \text{Km} \). The possibility of such an error deserves consideration because the \( \text{Km} \) was determined in the micromolar range of \( \text{H}_2\text{O}_2 \), whereas the concentration of \( \text{H}_2\text{O}_2 \) in these studies was largely in the nanomolar range. The agreement between expected and observed results (Fig 4), however, refutes that possibility. Earlier attention to the \( \text{Km} \) of glutathione peroxidase and to the activities of the two enzymes in erythrocytes would have allowed anticipation of the growing number of studies\(^{1,2,14-18} \) discounting the notion that glutathione peroxidase is the major means by which \( \text{H}_2\text{O}_2 \) is removed in erythrocytes. Aebi and Suter\(^6 \) mentioned the possible use of this approach in 1974. Because the assignment of a maximum rate for glutathione peroxidase, relative to that of catalase, does not depend on knowledge of the \( \text{H}_2\text{O}_2 \) concentration, this approach could be applicable to other tissues and species. The ping-pong kinetics of glutathione peroxidase would require, however, that the concentration of GSH be known before the \( \text{Km} \) of glutathione peroxidase for \( \text{H}_2\text{O}_2 \) could be calculated.\(^{20} \) Pending the outcome of those studies, we suggest that the function of glutathione peroxidase may be the disposal of organic peroxides rather than the removal of \( \text{H}_2\text{O}_2 \).

Our earlier studies indicated that catalase accounts for about half of the \( \text{H}_2\text{O}_2 \) removed in human erythrocytes.\(^1,2 \) Glutathione peroxidase was presumed to account for the other half. The present study indicates, however, that glutathione peroxidase could account for only a small portion of the remaining half, as measured by the rate of oxidation of NADPH and GSH. This relationship is reflected in Fig 5, where the rate of 6PG generation can be seen to decrease less with intact cells than with cell-free solutions, as progressively higher catalase activities are considered. A question requiring addressing is the nature of this unknown portion of \( \text{H}_2\text{O}_2 \)-removing ability. The well-recognized pseudo-peroxidase activity of hemoglobin\(^21 \) may account for at least some of the missing portion. At the relatively low concentrations of components used in this study, some increase in rate of 6PG production is consistently seen when hemoglobin is present (Table 2, Figs 2, 4, and 5). Catalase oxidizes NADPH in the process of protecting itself from being inactivated by its own toxic substrate (\( \text{H}_2\text{O}_2 \)).\(^{13} \) It is unlikely, however, that this accounts for much of the missing portion, as catalase oxidizes NADPH at less than 7% of the rate at which it converts \( \text{H}_2\text{O}_2 \) to oxygen and water.\(^{13} \)

**REFERENCES**


18. Scott MD, Wagner TC, Chiu DTY: Decreased catalase activity is the underlying mechanism of oxidant susceptibility in glucose-6-phosphate dehydrogenase-deficient erythrocytes. Biochim Biophys Acta 1181:163, 1993


Predominant role of catalase in the disposal of hydrogen peroxide within human erythrocytes

GF Gaetani, AM Ferraris, M Rolfo, R Mangerini, S Arena and HN Kirkman